



Cambridge IGCSE™

CHEMISTRY

0620/22

Paper 2 Multiple Choice (Extended)

October/November 2023

45 minutes

You must answer on the multiple choice answer sheet.

You will need: Multiple choice answer sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

INSTRUCTIONS

- There are **forty** questions on this paper. Answer **all** questions.
- For each question there are four possible answers **A**, **B**, **C** and **D**. Choose the **one** you consider correct and record your choice in soft pencil on the multiple choice answer sheet.
- Follow the instructions on the multiple choice answer sheet.
- Write in soft pencil.
- Write your name, centre number and candidate number on the multiple choice answer sheet in the spaces provided unless this has been done for you.
- Do **not** use correction fluid.
- Do **not** write on any bar codes.
- You may use a calculator.

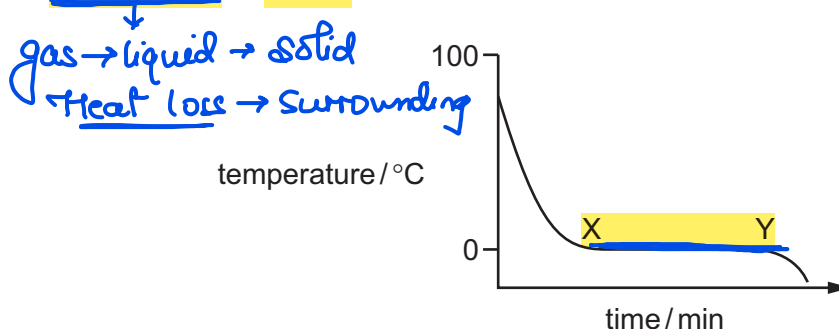
INFORMATION

- The total mark for this paper is 40.
- Each correct answer will score one mark.
- Any rough working should be done on this question paper.
- The Periodic Table is printed in the question paper.

This document has **16** pages.



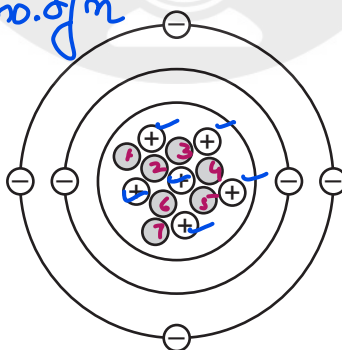
- 1 Part of a **cooling curve** for **water** is shown.



What is **occurring** between points X and Y?

- A** Steam is condensing into water.
B The temperature of the water is decreasing.
C Ice is melting.
D Particles are losing heat to the surroundings.
- 2 Which **statements** about **clean, dry air** are **correct**?
- 1 It is a **mixture of elements only**.
 - 2 It is a **mixture of elements and compounds**.
 - 3 It contains only non-metals.
- A** 1 and 3 **B** 1 only **C** 2 and 3 **D** 2 only
- 3 A **representation of an atom** is **shown**.

$$\begin{aligned} \text{nucleon no} &= \text{no. of } p + \text{no. of } n \\ &= 6p + 7n \\ &= \underline{\underline{13}} \end{aligned}$$



What is the **nucleon number** of this atom?

- A** 6 **B** 7 **C** 12 **D** 13



- 4 The percentage abundances of three isotopes in a sample of neon are shown.

isotope	percentage abundance / %
$^{20}_{10}\text{Ne}$	90.48
$^{21}_{10}\text{Ne}$	0.27
$^{22}_{10}\text{Ne}$	9.25

$$20 \times 90.48 + 21 \times 0.27 + 22 \times 9.25 = 2017.3$$

$$\frac{2017.3}{100} = 20.17$$

What is the relative atomic mass, A_r , of this sample of neon?

- A 10.19 **B 20.19** C 21.00 D 30.19

- 5 Potassium reacts with iodine to form potassium iodide.



Which statement about potassium iodide is correct?

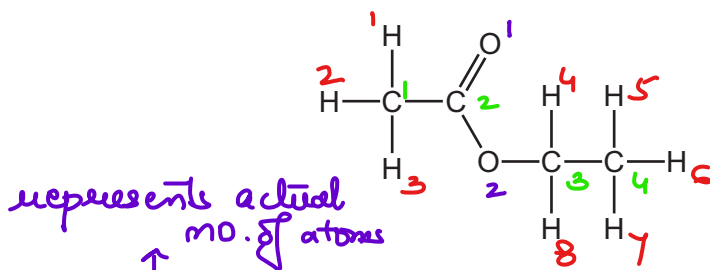


- ☒ A Each potassium atom shares a pair of electrons with an iodine atom.
- ☒ B In potassium iodide, the particles of potassium have more protons than electrons.
- ☒ C Potassium iodide has a high melting point because it is a covalent compound.
- ☒ D Potassium iodide has a low melting point because it is an ionic compound.

- 6 Which substance has the lowest melting point?

- A graphite → giant covalent structure
- ☒ B methanol CH_3OH
- C silicon(IV) oxide → giant structure
- D sodium chloride → ionic compound

- 7 The diagram shows the structure of a molecule of ethyl ethanoate.



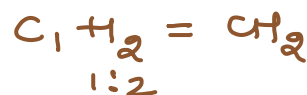
What is the molecular formula of a molecule of ethyl ethanoate?

- A CHO **B $C_4H_8O_2$** C $C_4(H_2)_2(O_2)$ D C_2H_4O

hydrogen & carbon
↑

4

C 85.7 85.7/12 7.14 7.14/14 1
H 14.3 14.3/1 14.3 14.3/7.14 2



- 8 A hydrocarbon contains 85.7% of carbon by mass.

What is the empirical formula of the hydrocarbon?

- A ☒ CH₂ B CH₄ C C₂H₅ D C₃H₆

- 9 The formula of a compound containing element X is Na₂X₂O₃.

The relative formula mass of the compound is 158.

What is the relative atomic mass of X?

- A ☒ 32 B 59.5 C 64 D 119

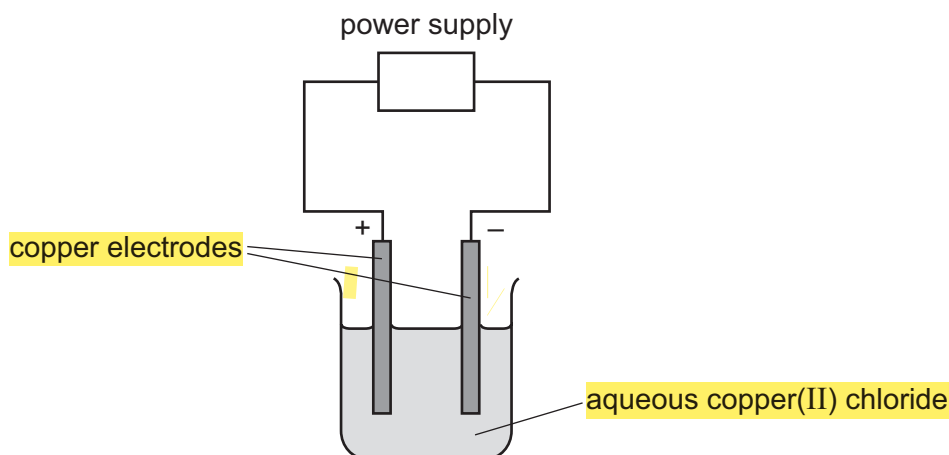
$$\begin{aligned} 2 \times 23 + 2x + 3 \times 16 &= 158 \\ 46 + 2x + 48 &= 158 \\ 2x + 94 &= 158 \\ 2x &= 158 - 94 \\ 2x &= 64 \\ x &= \frac{64}{2} = 32 \end{aligned}$$

- 10 Dilute aqueous potassium chloride is electrolysed using platinum electrodes.

Which row identifies the product at each electrode?

	+ve anode	-ve cathode
A	chlorine	hydrogen
B	chlorine	potassium
☒ C	oxygen	hydrogen
D	oxygen	potassium

- 11 Concentrated aqueous copper(II) chloride is electrolysed using copper electrodes, as shown.



What happens to the mass of each electrode during this process?

	positive electrode	negative electrode
A	decreases	decreases
☒ B	decreases	increases
C	increases	decreases
D	increases	increases

- 12 The initial and final temperatures of four different reactions are measured.

Which reaction is the least exothermic? *→ difference should be less*

	initial temperature / °C	final temperature / °C	
A	19	25	<i>→ 25 - 19</i>
B	21	18	<i>→ 18 - 21</i>
C	22	17	<i>→ 17 - 22</i>
☒ D	22	26	<i>→ 26 - 22</i>

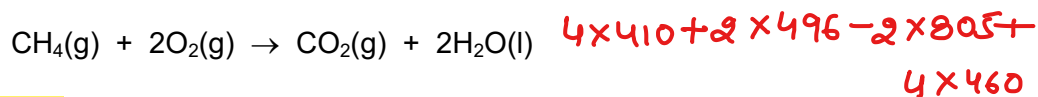
Heat absorbed

- 13 Which equation represents an endothermic reaction?

- ☒ A $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$ *cl { cl*
- B $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ *→ combustion rxn*
- C $\text{H}(\text{g}) + \text{H}(\text{g}) \rightarrow \text{H}_2(\text{g})$ *→ combination*
- D $2\text{K}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{KOH}(\text{aq}) + \text{H}_2(\text{g})$



14 Methane burns in oxygen to form carbon dioxide and water.



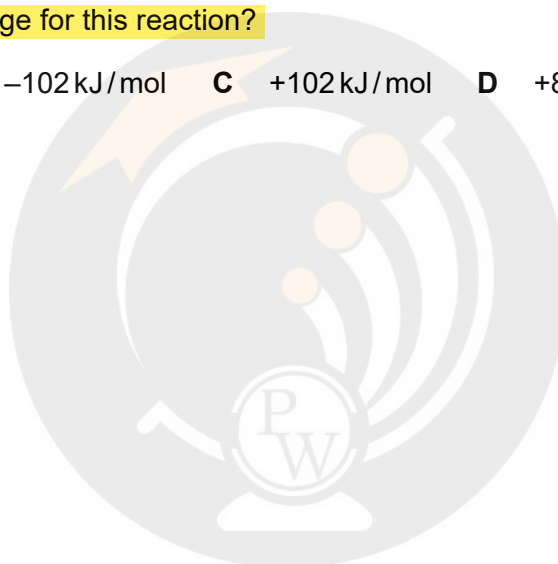
The bond energies are shown.

bond	bond energy in kJ/mol
C-H	410
C-O	360
C=O	805
O-H	460
O-O	146
O=O	496

$$= 1640 + 992 - 1610 + 1840$$
$$= -818$$

What is the energy change for this reaction?

- ☒ A -818 kJ/mol B -102 kJ/mol C +102 kJ/mol D +818 kJ/mol

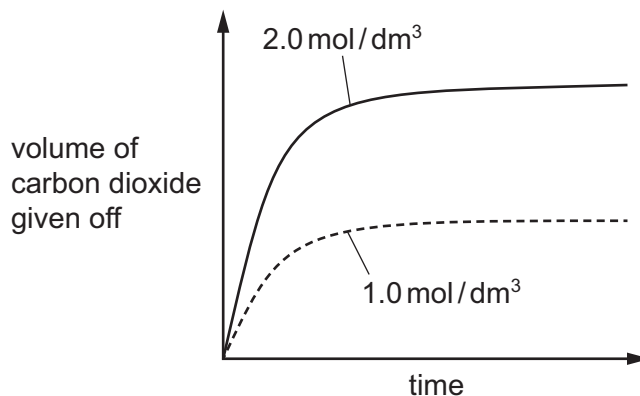




- 15 Hydrochloric acid is added to excess calcium carbonate in two separate experiments.

Two different concentrations of hydrochloric acid are used but the temperature is the same in both experiments.

The graph of the results shows the volume of carbon dioxide gas given off over time.

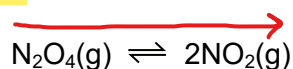


Which row is correct?

	particles in 2.0 mol/dm ³ compared to 1.0 mol/dm ³	
	collision rate	collision energy
A	higher	no change
B	higher	higher
C	lower	no change
D	lower	higher

- 16 The decomposition of dinitrogen tetroxide, N₂O₄, into nitrogen dioxide, NO₂, is a reversible reaction.

The equation for the reaction is shown.



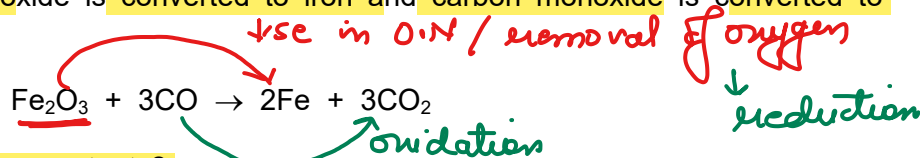
The forward reaction is endothermic.

Which row shows the effect on the position of equilibrium and the rate of the reverse reaction when the temperature is increased?

	position of equilibrium	rate of the reverse reaction
A	shifts to the left	decreases
B	shifts to the left	increases
C	shifts to the right	decreases
D	shifts to the right	increases



- 17 In a blast furnace, iron(III) oxide is converted to iron and carbon monoxide is converted to carbon dioxide.



What happens to each of these reactants?

- A Both iron(III) oxide and carbon monoxide are oxidised.
- B Both iron(III) oxide and carbon monoxide are reduced.
- C Iron(III) oxide is oxidised and carbon monoxide is reduced.
- ☒ D Iron(III) oxide is reduced and carbon monoxide is oxidised.

- 18 Which row describes what happens to Fe^{2+} ions when they are oxidised?

	electron movement	oxidation number of iron
A	they gain electrons	decreases
B	they gain electrons	increases
C	they lose electrons	decreases
☒ D	they lose electrons	increases

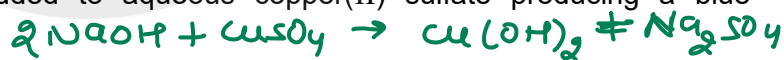
- 19 In which reaction does an acid react with a base?



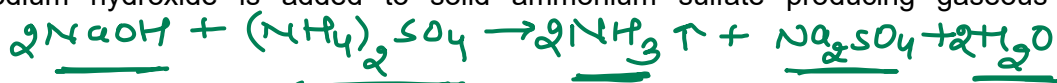
- A Dilute sulfuric acid is added to a piece of magnesium ribbon producing hydrogen.
- B Dilute sulfuric acid is added to aqueous barium chloride producing a white precipitate of barium sulfate.



- C Aqueous sodium hydroxide is added to aqueous copper(II) sulfate producing a blue precipitate of copper(II) hydroxide.



- ☒ D Aqueous sodium hydroxide is added to solid ammonium sulfate producing gaseous ammonia.



- 20 Which element forms an oxide that reacts with an aqueous solution of a base?

- A argon → inert

acidic

- ☒ B sulfur → non-metal → acidic oxide

- C magnesium } metals → basic oxide

- D copper



9

→ precipitate

21 Which method is used to produce insoluble salts?

- A addition of excess insoluble base to an acid
- B addition of excess metal to an acid
- ☒ C precipitation using two aqueous solutions
- D titration using an acid and an alkali

22 The noble gases are in Group VIII of the Periodic Table.

Some properties of the first four noble gases are shown.

noble gas	boiling point in °C	density in g/dm ³
helium	-267	0.179
neon	-246	0.900
argon	-186	1.782
krypton	-152	3.708

Which row identifies the trends in boiling point and in density as Group VIII is descended?

	boiling point	density
A	decreasing	increasing
☒ B	increasing	increasing
C	decreasing	decreasing
D	increasing	decreasing

23 Some properties of element R are shown.

melting point in °C	98
boiling point in °C	883
reaction with cold water	gives off H ₂ gas
reaction when heated with oxygen	burns to give a white solid

In which part of the Periodic Table is R found?

- ☒ A Group I
- B Group VII → halogen
- C Group VIII → inert gas
- D transition elements → high m.pt



24 Which pair of compounds shows that transition elements have variable oxidation states?

- A Cr_2O_3 and CrBr_3
B CuSO_4 and CuCl_2
✓ C Fe_2O_3 and FeCl_2
D NiO and NiCl_2

25 The list gives the order of some metals and hydrogen in the reactivity series.

Metal X is also included.

most reactive K
 Mg
 Zn
 H
 X
least reactive Cu

Which row shows the properties of metal X?

	reacts with dilute acids	oxide reduced by carbon
A	no	no
✓ B	no	yes
C	yes	no
D	yes	yes

26 When zinc is added to an aqueous solution containing ^{Group 2} magnesium ions, there is no reaction.

Which species has the greatest tendency to ^{Zn} lose electrons?

- ✓ A Mg B Mg^{2+} C Zn D Zn^{2+}

27 Which gas in the air is needed for iron to rust?

- A argon
B carbon dioxide
C nitrogen
✓ D oxygen

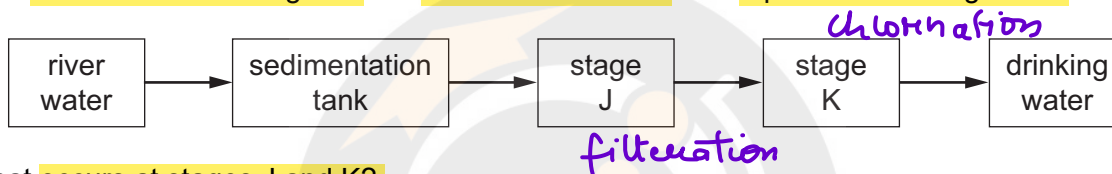
28 Which coating prevents iron from rusting even when the coating is damaged?

- A grease
- B paint
- C plastic
- ☒ D zinc

29 Why is limestone added to the blast furnace?

- A It neutralises the molten slag produced.
- ☒ B It reacts with impurities to form slag.
- C It releases carbon dioxide which reduces the iron(III) oxide.
- D It removes acidic gases such as carbon dioxide.

30 The flow chart shows stages in the treatment of river water to produce drinking water.



What occurs at stages J and K?

	J	K
A	distillation	chlorination
B	distillation	filtration
☒ C	filtration	chlorination
D	filtration	distillation

- 31 Carbon dioxide acts as a greenhouse gas by interacting with a particular type of energy that radiates from the Earth's surface into the atmosphere.

Which type of energy is involved and what happens when this energy interacts with carbon dioxide molecules?

	type of energy involved	what happens
A	thermal	carbon dioxide molecules increase the Earth's energy loss to space
☒ B	thermal	carbon dioxide molecules absorb the energy
C	light	carbon dioxide molecules increase the Earth's energy loss to space
D	light	carbon dioxide molecules absorb the energy

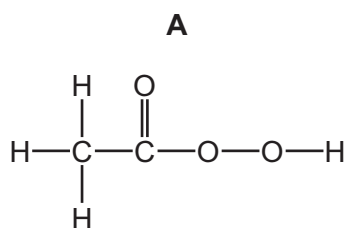
- 32 Oxides of nitrogen, such as NO and NO₂, are formed in the petrol engines of cars.

They are removed from the exhaust gases by reactions in the car's catalytic converter.

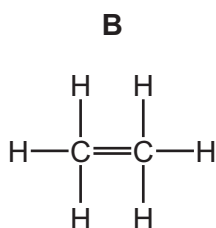
Which row describes how oxides of nitrogen are formed in a petrol engine and a reaction that happens in the catalytic converter?

	how oxides of nitrogen are formed	a reaction that happens in the catalytic converter
☒ A	by the reaction between nitrogen and oxygen from the air	$2\text{NO} + 2\text{CO} \rightarrow \text{N}_2 + 2\text{CO}_2$
B	by the reaction between nitrogen and oxygen from the air	$2\text{NO} + 2\text{H}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$
C	by the reaction between nitrogen compounds in petrol and oxygen from the air	$2\text{NO} + 2\text{CO} \rightarrow \text{N}_2 + 2\text{CO}_2$
D	by the reaction between nitrogen compounds in petrol and oxygen from the air	$2\text{NO} + 2\text{H}_2 \rightarrow \text{N}_2 + 2\text{H}_2\text{O}$

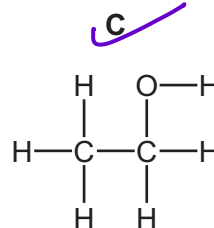
- 33 Which diagram shows the displayed formula for the named organic compound?



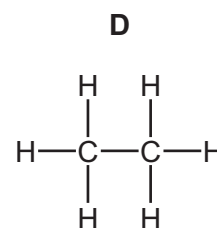
ethanoic acid



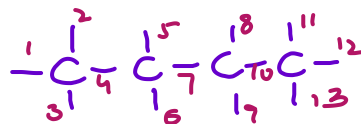
ethene



ethanol



methane



34 What is the total number of covalent bonds in a molecule of butane, C_4H_{10} ?

A 3

B 10

C 13

D 14

alkane

35 Propane reacts with chlorine in a substitution reaction.

Which reaction condition is required for the reaction to occur?

A acid catalyst

B iron catalyst

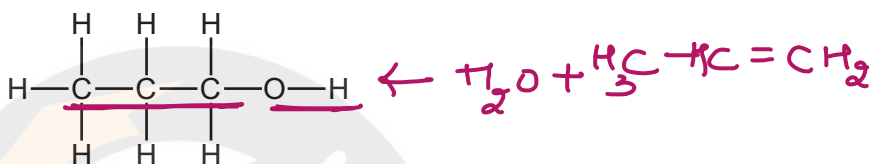
C temperature of 400°C

D ultraviolet light



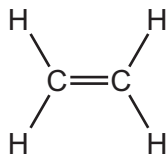
36 The structure of an organic compound is shown.

*addition reaction
is characteristic
reaction for alkene*

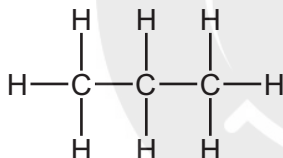


Which structure represents a molecule that reacts with steam to produce this product?

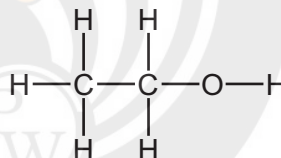
A



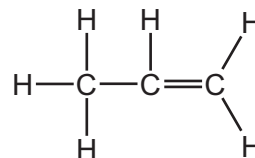
B



C



D



37 Which term describes nylon?

A addition polymer

B natural polymer

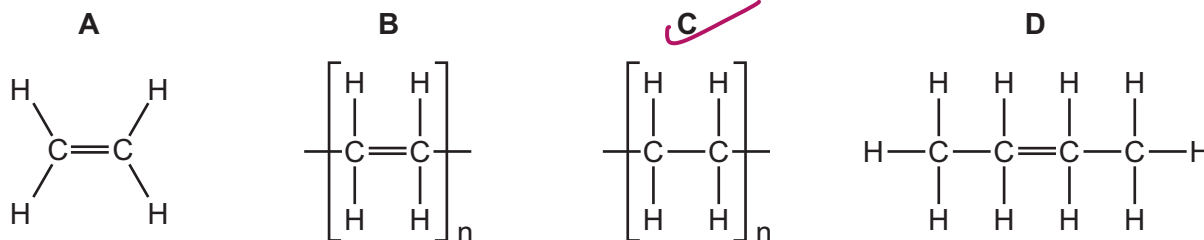
C polyamide

D polyester

*hexamethylene diamine
adipic acid*

38 Ethene can be polymerised.

Which diagram represents the structure of the product formed?



39 An acid–base titration is described.

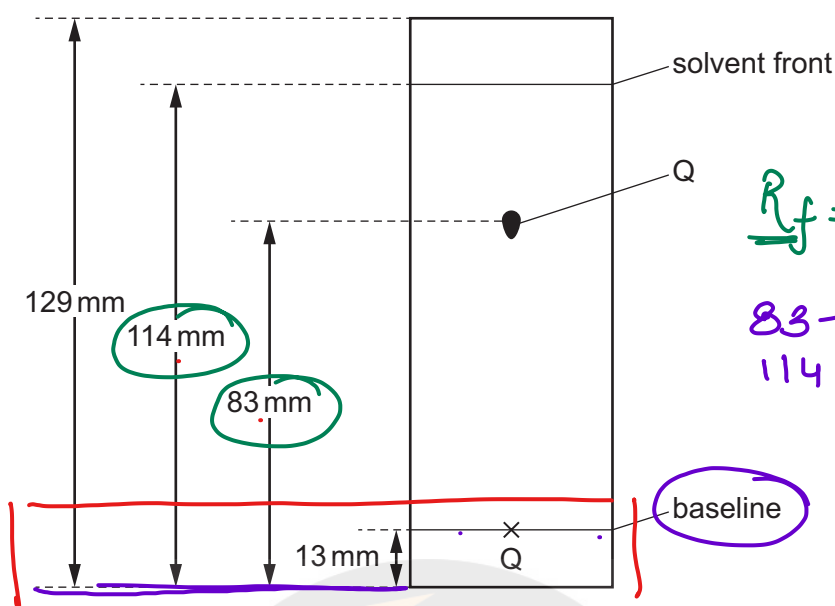
- 25.0 cm³ of dilute aqueous alkali is put into a conical flask.
- Indicator is added to the flask.
- Dilute acid is added to the aqueous alkali until the indicator changes colour.
- The volume of acid used is then recorded.

Which use of apparatus is correct?

- A**  The 25.0 cm³ of aqueous alkali is measured using a volumetric pipette.
- ~~B~~  The 25.0 cm³ of aqueous alkali is measured using the lines on the conical flask.
- ~~C~~  The volume of acid is measured using a measuring cylinder.
- ~~D~~  The volume of acid is measured using a volumetric pipette.

40 Substance Q is investigated using chromatography.

The chromatogram is shown. The diagram is not drawn to scale.



What is the R_f value of Q?

A 0.60

B 0.64

☒ C 0.69

D 0.72

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The Periodic Table of Elements

Group																							
I	II	1 Hhydrogen1												III	IV	V	VI	VII	VIII				
		Key atomic number atomic symbol name relative atomic mass																					
3	4																	5	6	7	8	9	
Li	Be																	B	C	N	O	F	
lithium	beryllium																	boron	carbon	nitrogen	oxygen	fluorine	
7	9																	11	12	14	16	17	
Na	Mg																	Al	Si	P	S	Cl	
sodium	magnesium																	aluminium	silicon	phosphorus	sulfur	chlorine	
23	24																	27	28	31	32	35.5	
19	20																	31	32	33	34	35	
K	Ca																	Ga	Ge	As	Se	Br	
potassium	calcium																	gallium	germanium	arsenic	selenium	bromine	
39	40																	70	73	75	79	80	
37	38																	49	50	51	52	53	
Rb	Sr																	In	Sn	Sb	Te	I	
rubidium	strontium																	indium	tin	antimony	tellurium	iodine	
85	88																	115	119	122	128	127	
55	56																	81	82	83	84	85	
Cs	Ba																	Tl	Pb	Bi	Po	At	
caesium	barium																	thallium	lead	bismuth	polonium	astatine	
133	137																	204	207	209	207	—	
87	88																	113	114	115	116	117	
Fr	Ra																	Nh	Fl	Mc	Lv	Ts	
francium	radium																	nihonium	flerovium	moscovium	livermorium	tennessine	
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lanthanoids

actinoids

57	58	59	60	61	62	63	64	65	66	67	68	69	70	71
La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
lanthanum	cerium	praseodymium	neodymium	promethium	samarium	europium	gadolinium	terbium	dysprosium	holmium	erbium	thulium	ytterbium	lutetium
139	140	141	144	—	150	152	157	159	163	165	167	169	173	175
89	90	91	92	93	94	95	96	97	98	99	100	101	102	103
Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr
actinium	thorium	protactinium	uranium	neptunium	plutonium	americium	curium	berkelium	californium	einsteinium	fermium	mendelevium	nobelium	lawrencium
—	232	231	238	—	—	—	—	—	—	—	—	—	—	—

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).