

JEE MAIN 2025

PAPER DISCUSSION

Attempt : 01

Date : 28th Jan 2025

Shift : 01



Inorganic *Chemistry*

The molecules having square pyramidal geometry are



A ~~SbF₅ and PCl₅~~

B ~~BrF₅ and XeOF₄~~

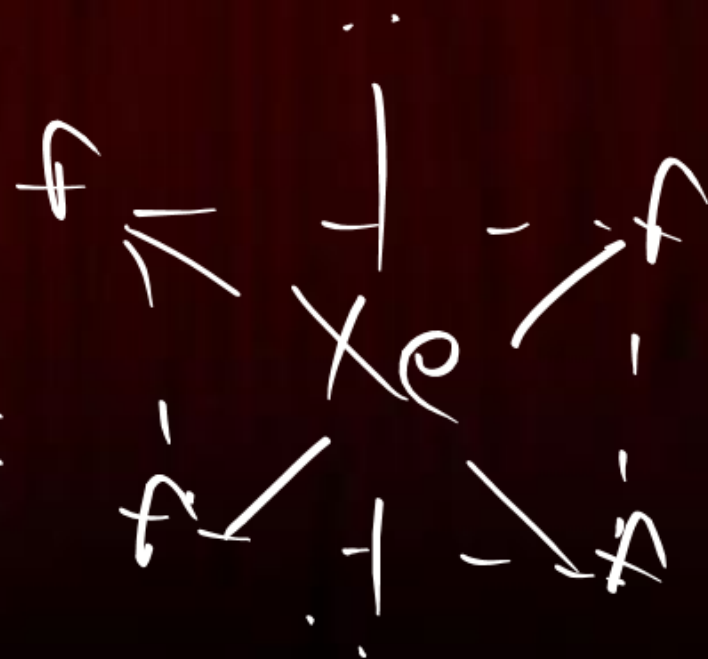
C ~~BrF₅ and PCl₅~~

D ~~SbF₅ and XeF₄~~

$lp \rightarrow 0, bp \rightarrow 5, ep \rightarrow 5, \underline{TBP}$

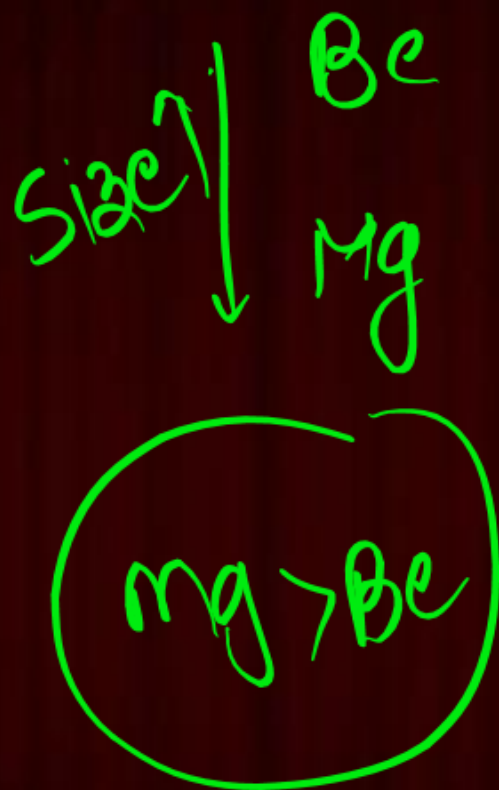
$lp \rightarrow 1, bp \rightarrow 5, ep \rightarrow 6$

$bp \rightarrow 4, lp \rightarrow 2, ep \rightarrow 6$



The incorrect decreasing order of atomic radii is:

- A $\text{Si} > \text{P} > \text{Cl} > \text{F}$
- B $\text{Mg} > \text{Al} > \text{C} > \text{O}$
- C $\text{Al} > \text{B} > \text{N} > \text{F}$
- ~~D $\text{Be} > \text{Mg} > \text{Al} > \text{Si}$~~



Consider the following elements; In, Tl, Al and Pb. The most stable oxidation states of elements with highest and lowest first ionisation enthalpies, respectively are:

A +4 and +1

B +2 and +3

C +4 and +3

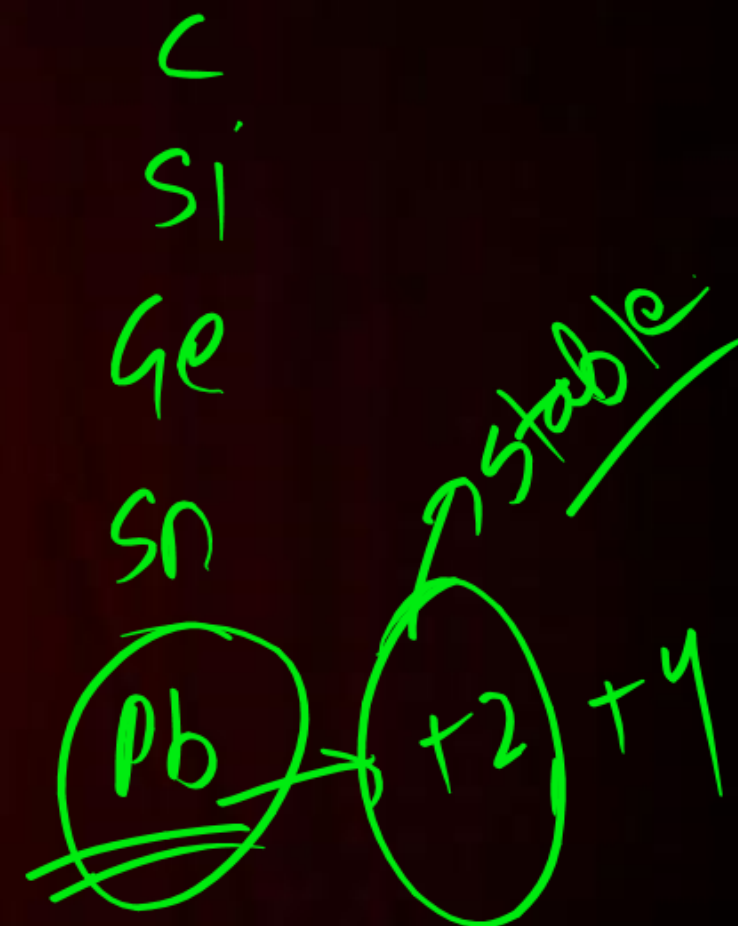
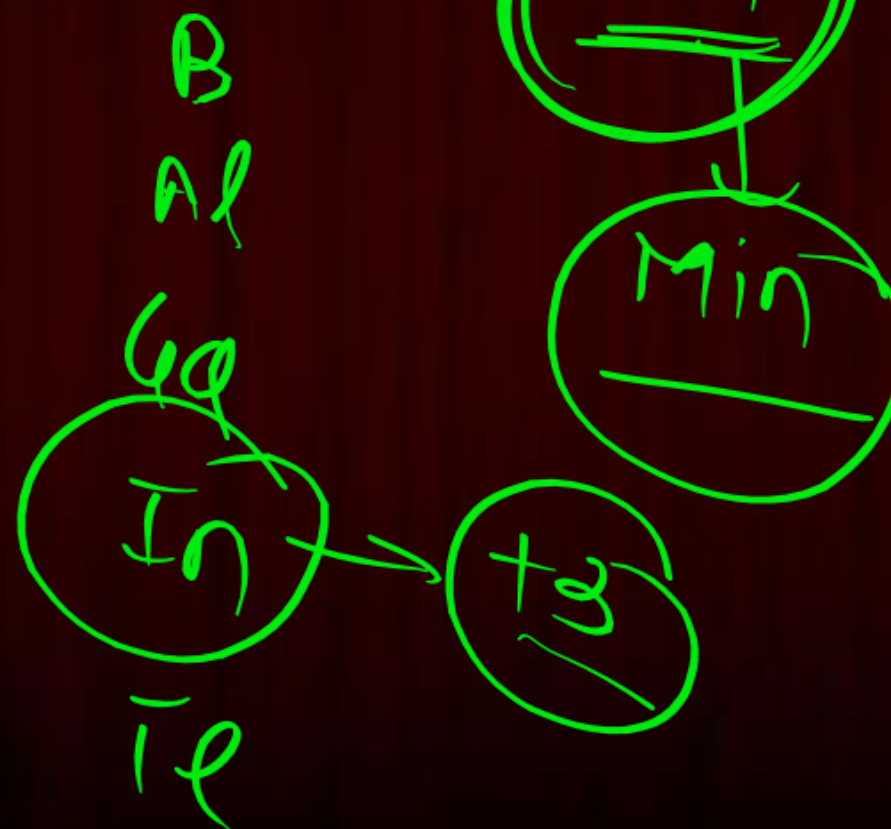
D +1 and +4

+2

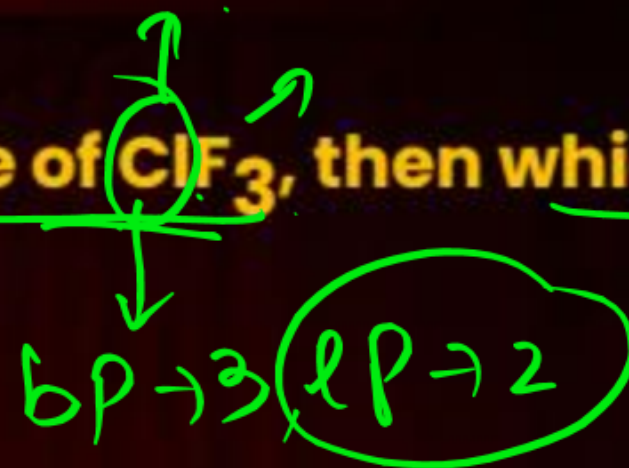
G-13

I.E. $\rightarrow$ $G-13 < G-14$

B >



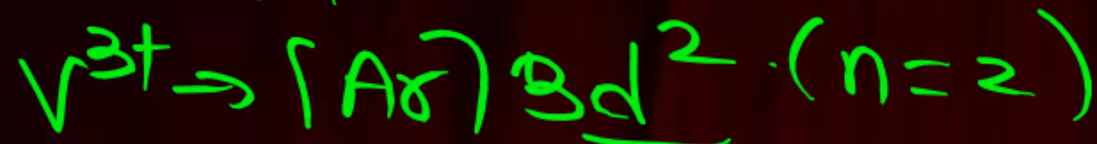
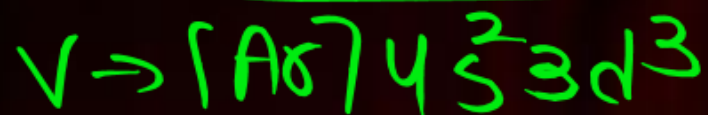
Let n be the number of lone pairs for most stable structure of ClF₃, then which of the following has n number of unpaired electrons?



n = 2

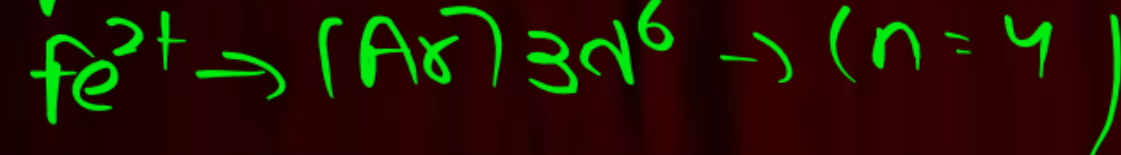
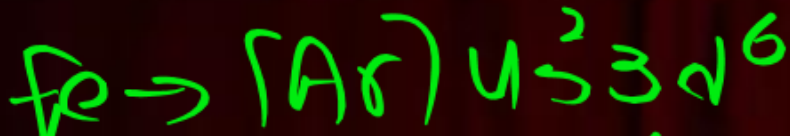
☒ A

V³⁺



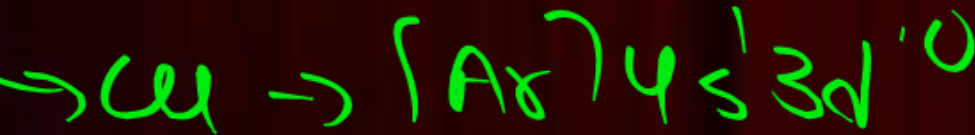
☒ B

Fe²⁺



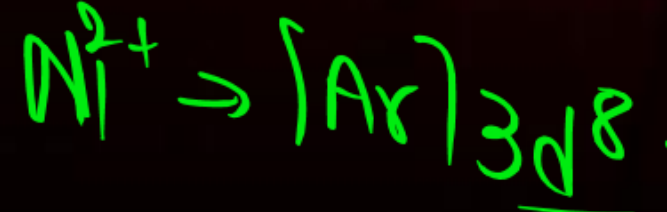
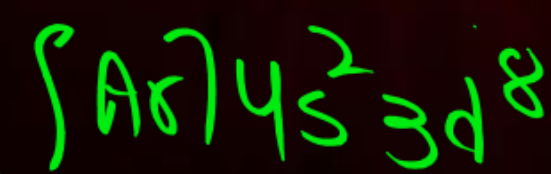
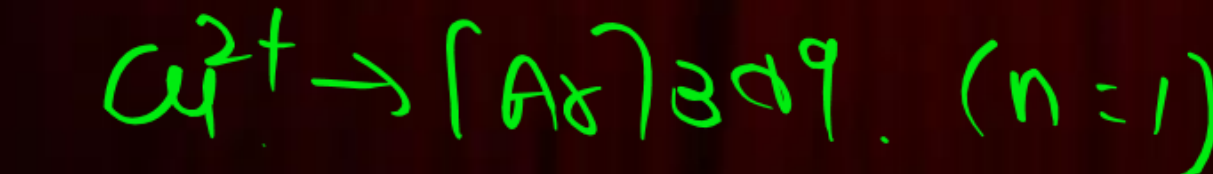
☒ C

Cu²⁺

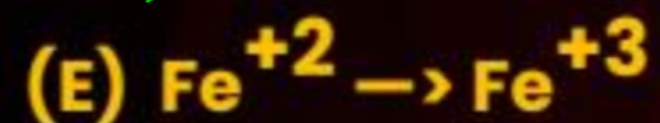
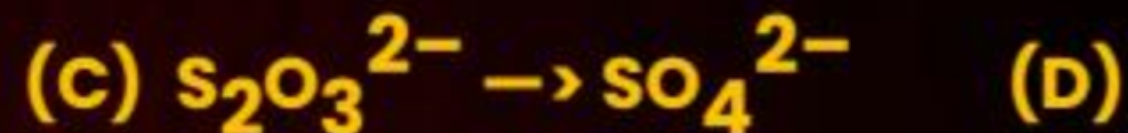
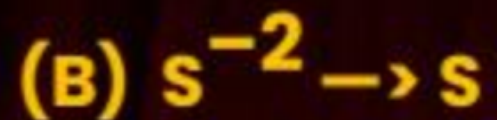
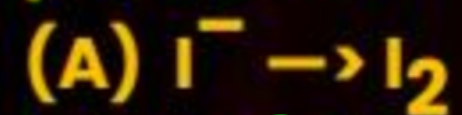


☒ D

Ni²⁺



On reaction with $\text{K}_2\text{Cr}_2\text{O}_7$ or KMnO_4 in acidic medium. How many reactions are possible?



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