

PUC PYQ 2023

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CHEMISTRY

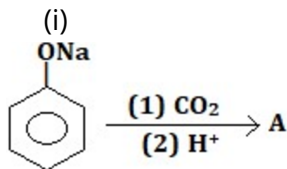
- Q1** Which kind of defect is introduced by doping intrinsic semiconductors? [PUC 2023]
 (A) Dislocation defect
 (B) Electronic defect
 (C) Interstitial defect
 (D) Schottky defect
- Q2** A binary liquid mixture that forms maximum boiling azeotrope at a specific composition is [PUC 2023]
 (A) Ethanol + Water
 (B) n - Hexane + n- Heptane
 (C) Benzene + Toluene
 (D) Nitric acid + Water
- Q3** The value of van't Hoff factor (i) for ethanoic acid in benzene is nearly [PUC 2023]
 (A) 2 (B) 1
 (C) 0.5 (D) 0
- Q4** On charging the Lead storage battery, PbSO_4 on cathode is converted into [PUC 2023]
 (A) PbO_2 (B) Pb
 (C) PbO (D) PbO
- Q5** In the Arrhenius equation the factor $e^{-\frac{E_a}{RT}}$ corresponds to [PUC 2023]
 (A) Collision frequency
 (B) Proper orientation
 (C) The fraction of molecules with kinetic energy $> E_a$
 (D) Threshold energy
- Q6** Which one of the following is not applicable to the phenomenon of adsorption? [PUC 2023]
 (A) $\Delta G = -Ve$
 (B) $\Delta S = -Ve$
 (C) $\Delta H = -Ve$
 (D) $\Delta H = +Ve$
- Q7** What is the role of NaCN in the separation of ZnS and PbS by froth floatation method ? [PUC 2023]
 (A) depressant
 (B) froth stabilizer
 (C) collector
 (D) reductant
- Q8** On complete hydrolysis of XeF_6 with water, the product formed is [PUC 2023]
 (A) XeF_4 (B) XeO_3
 (C) XeO_2F_2 (D) XeOF_4
- Q9** Which of the following element is not regarded as transition element? [PUC 2023]
 (A) Fe (B) Mn
 (C) Sc (D) Zn
- Q10** M C- bond in metal carbonyls possesses _____. [PUC 2023]
 (A) Ionic character
 (B) Both σ and π characters
 (C) π -character only
 (D) Ion-dipole forces
- Q11** Identify chiral molecule in the following compounds. [PUC 2023]
 (A) 2- Bromobutane
 (B) 1-Bromobutane
 (C) 2- Bromopropane
 (D) 2-Bromo-2-methyl-Propane



- Q12** When CH_3ONa reacts with $(\text{CH}_3)_3\text{CBr}$, it gives exclusively [PUC 2023]
 (A) t- Butylmethyl ether
 (B) 2,2-Dimethyl propane
 (C) 2- Methyl propene
 (D) 2-Methyl Propan -2-ol
- Q13** Iodoform reaction with NaOI can be used for the detection of the compound [PUC 2023]
 (A) $\text{C}_2\text{H}_5\text{COC}_2\text{H}_5$ (B) CH_3CHO
 (C) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ (D) $(\text{CH}_3)_3\text{COH}$
- Q14** Nitration of aniline in the strongly acidic medium at 288 K yields [PUC 2023]
 (A) 2,4,6 – Trinitroaniline
 (B) o and p – Nitroanilines
 (C) m- Nitroaniline
 (D) o, m, and p – Nitroanilines
- Q15** Which hormone is an iodinated derivative of amino acid tyrosine? [PUC 2023]
 (A) Insulin (B) Epinephrine
 (C) Thyroxin (D) Glucagon
- Q16** Because of low concentration of O_2 in the blood and tissues of people living at high altitudes, suffer from a disease called _____ [PUC 2023]
- Q17** Inversion of cane sugar is an example of _____ reaction. [PUC 2023]
- Q18** Radon is obtained as a decay product of _____. [PUC 2023]
- Q19** When Chlorobenzene is treated with sodium in dry ether _____ is formed. [PUC 2023]
- Q20** _____ is a synthetic progesterone derivative, most widely used as an antifertility drug. [PUC 2023]
- Q21** Give any two differences between Frenkel defect and Schottky defect. [PUC 2023]
- Q22** Λ_m^0 for NaCl , HCl and NaAc (Sodium acetate) are $126.4 \text{ S cm}^2 \text{ mol}^{-1}$, $425.9 \text{ S cm}^2 \text{ mol}^{-1}$ and $91.0 \text{ S cm}^2 \text{ mol}^{-1}$ respectively. Calculate Λ_m^0 for HAc (acetic acid). [PUC 2023]
- Q23** What are the two criteria for the effective collisions between molecules in a chemical reaction ? [PUC 2023]
- Q24** Give reason:
 (a) Actinoids exhibit a greater range of oxidation states.
 (b) Zr and Hf have the almost identical atomic radii. [PUC 2023]
- Q25** What happens when Phenol is heated with Zinc dust? Write equation. [PUC 2023]
- Q26** How is Benzoyl chloride converted into Benzaldehyde? Name the reaction. [PUC 2023]
- Q27** What is the role of following chemicals in food?
 (a) Sodium benzoate
 (b) Butylated Hydroxy Anisole (BHA). [PUC 2023]
- Q28** Why soap does not work in hard water? [PUC 2023]
- Q29** Explain the extraction of 'blister copper' from copper matte. Write the balanced equations for the reactions taking place in then convertor. [PUC 2023]
- Q30** Write the chemical equations with reaction conditions involved in the manufacture of Nitric acid by Ostwald's process. [PUC 2023]
- Q31** Complete the following equation:
 (a) $\text{PbS} + 4\text{O}_3 \rightarrow \text{_____} + 4\text{O}_2$
 (b) $5\text{SO}_2 + 2\text{MnO}_4^- + 2\text{H}_2\text{O} \rightarrow 5\text{SO}_4^{2-} + 4\text{H}^+ + \text{_____}$
 (c) $\text{C}_{12}\text{H}_{22}\text{O}_{11} \xrightarrow{\text{conc. H}_2\text{SO}_4} \text{_____} + 11\text{H}_2\text{O}$ [PUC 2023]



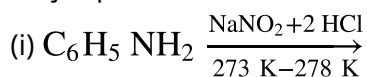
- Q32** (a) How is chlorine manufactured by Deacon's process? Give equation.
(b) Write the structure of Chlorous acid. [PUC 2023]
- Q33** (a) The transition metals and their compounds are known for their catalytic activity.
Give two reasons.
(b) What is Mischmetall? [PUC 2023]
- Q34** Explain the preparation of Potassium permanganate from MnO_2 with equations. [PUC 2023]
- Q35** Out of the following two coordination entities; $\text{cis} - [\text{PtCl}_2(\text{en})_2]^{2+}$ and $\text{cis} - [\text{PtCl}_2(\text{en})_2]^{2+}$.
(a) Which is Chiral (optically active)?
(b) Draw the structures of its enantiomers. [PUC 2023]
- Q36** According to Valence Bond Theory [VBT], explain geometry, hybridisation and magnetic property of $[\text{CoF}_6]^{-3}$ ion. [Atomic number of Cobalt is 27]. [PUC 2023]
- Q37** Calculate the packing efficiency in Face centred Cubic (FCC) Lattice.
Potassium metal crystallises in a bcc unit cell with edge length 542 pm. Calculate the density of potassium metal.
(Atomic mass of K = 39 g mol^{-1} , $N_A = 6.022 \times 10^{23} \text{ atoms mol}^{-1}$). [PUC 2023]
- Q38** (a) 450 cm^3 of an aqueous solution of a protein contains 1.0 g of the protein.
The osmotic pressure of such a solution at 310 K is found to be $3.1 \times 10^{-4} \text{ bar}$.
Calculate the molar mass of the protein. ($R = 0.083 \text{ L bar mol}^{-1} \text{ K}^{-1}$).
(b) State Raoult's law of relative lowering of vapour pressure. Write its mathematical form. [PUC 2023]

- Q39** (a) Calculate the standard Gibb's energy ($\Delta_r G^0$) for the reaction at 298 K :
 $\text{Zn}_{(s)} + 2 \text{Ag}_{(aq)}^+ \rightarrow \text{Ag}_{(s)} + \text{Zn}_{(aq)}^{2+}$
[Given : $E_{(\text{Zn}^{2+}/\text{Zn})} = -0.76 \text{ V}$ and $E_{(\text{Ag}^+/\text{Ag})} = +0.80 \text{ V}$; and $F = 96,500 \text{ C mol}^{-1}$]
(b) Write the balanced equations for the reactions taking place at anode and cathode during rusting of iron. [PUC 2023]
- Q40** (a) Derive an integrated rate equation for the rate constant of a first order reaction.
(b) Draw a graph of potential energy v/s reaction coordinate showing the effect of catalyst on the rate of a reaction. [PUC 2023]
- Q41** (a) Explain Bredig's Arc method for the preparation of metal sols.
(b) Write two steps involved in the mechanism of enzyme catalysed reaction. [PUC 2023]
- Q42** (a) Explain S_N1 mechanism of conversion of tert-butyl bromide to tert-butyl alcohol.
(b) Give any two reasons for the less reactivity of aryl halides towards nucleophilic substitution reactions. [PUC 2023]
- Q43** (a) Write the mechanism of the following reaction:
(b)
- $$\text{C}_2\text{H}_5\text{OH} \xrightarrow[443 \text{ K}]{\text{Conc. H}_2\text{SO}_4} \text{CH}_2 = \text{CH}_2 + \text{H}_2\text{O}$$
- (i)
- 
- (ii) $(\text{CH}_3)_3\text{C} - \text{OC}_2\text{H}_5 \xrightarrow{\text{HI}} \text{C}_2\text{H}_5\text{OH} + \text{B}$
[PUC 2023]

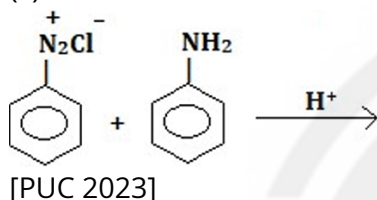


- Q44** (a) How is ketone prepared from Grignard reagent and nitrile? Explain with an example.
 (b) Explain Hell – Volhard – Zelinsky reaction. Give equation.
 (c) What is the role of dry HCl gas in the addition of alcohols to aldehydes? [PUC 2023]

- Q45** (a) Write the equations of reactions involved in the Gabriel Phthalimide synthesis of a primary amine.
 (b) Complete the following reactions by giving major products



(ii)



- Q46** (a) Write the Haworth's structure of lactose.
 (b) What is denaturation of proteins? Which level of structure remains intact during denaturation of globular proteins?
 (c) Name the sugar moiety present in DNA. [PUC 2023]

- Q47** (a) How is Buna – N prepared? Give equation.
 (b) Name the monomers of the biodegradable polymer Nylon -2-nylon-6.
 (c) Write the partial structure of Dacron. [PUC 2023]



Answer Key

Q1 B, C
Q2 D
Q3 C
Q4 A, B
Q5 C
Q6 D
Q7 A
Q8 B
Q9 D
Q10 B
Q11 A
Q12 A
Q13 A
Q14 A, B, C, D
Q15 C
Q16 (H & S)
Q17 (H & S)
Q18 (H & S)
Q19 (H & S)
Q20 (H & S)
Q21 (H & S)
Q22 (H & S)
Q23 (H & S)
Q24 (H & S)

Q25 (H & S)
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Q44 (H & S)
Q45 (H & S)
Q46 (H & S)
Q47 (H & S)



Hints & Solutions

Note: scan the QR code to watch video solution

Q1 Text Solution:

- (b) Electronic defect
OR
(c) Interstitial defect

Q2 Text Solution:

Specific composition 68% Nitric acid and 32% H_2O is not mentioned to get maximum boiling point azeotrope. For any option marks should be allotted

Q3 Text Solution:

0.5 or $\frac{1}{2}$

Q4 Text Solution:

- (a) PbO_2
OR
(b) Pb

Q5 Text Solution:

The fraction of molecules with kinetic energy $> E_a$

Q6 Text Solution:

$\Delta H = +Ve$

Q7 Text Solution:

Depressant

Q8 Text Solution:

XeO_3

Q9 Text Solution:

Zn

Q10 Text Solution:

Both σ and π characters

Q11 Text Solution:

The given particular options are not discussed under chiral concept in NCERT text book. For any option marks should be allotted

Q12 Text Solution:

Since it is higher application type of question marks should be allotted for any options.

Q13 Text Solution:

The given options are not discussed under iodoform reaction concept in NCERT text book. For any option marks should be allotted

Q14 Text Solution:

- (a) 2, 4, 6 – Trinitroaniline
(b) o and p – Nitroanilines
(c) m- Nitroaniline
(d) o, m, and p – Nitroanilines

Q15 Text Solution:

Thyroxin

Q16 Text Solution:

Anoxia

Q17 Text Solution:

Pseudo first order

Q18 Text Solution:

Radium -226

Q19 Text Solution:

Diphenyl

Q20 Text Solution:

Norethindrone

Q21 Text Solution:

Frenkel defect	Schottky defect
The smaller ion (cat ion) is dislocated from its normal site to an interstitial site.	It is caused by missing of equal number of cat ions and anions from lattice points to maintain electrical neutrality.
It is shown by ionic substances in which there is a large difference in the size of ions	It is shown by ionic substances in which the cation and anion are of almost similar sizes.
Density is not changed	Density decreases
It creates both vacancy and interstitial defects.	It creates vacancy defect.

Q22 Text Solution:

$$\begin{aligned}\Lambda_{\text{m(HAc)}}^0 &= \Lambda_{\text{m(HCl)}}^0 + \Lambda_{\text{m(NaAc)}}^0 - \Lambda_{\text{m(NaCl)}}^0 \\ &= 425.9 + 91.0 - 126.4 \\ &= 390.5 \text{ S cm}^2 \text{ mol}^{-1}\end{aligned}$$

Q23 Text Solution:

- (i) Activation energy OR Sufficient kinetic energy.
(ii) Proper orientation of molecules.

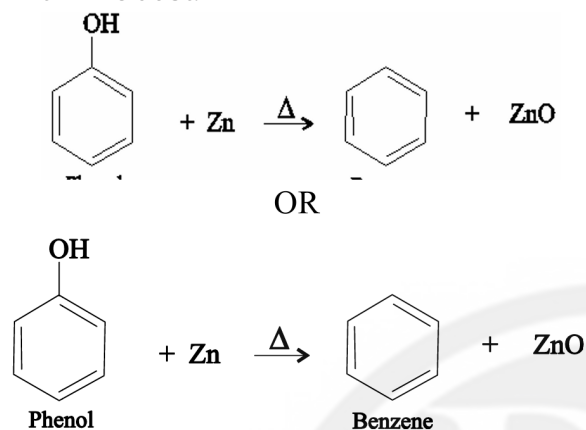


Q24 Text Solution:

- (a) Due to very comparable energies (small energy gap) between 5f, 6d and 7s subshells.
 (b) Due to lanthanoid contraction

Q25 Text Solution:

Phenol is converted to benzene on heating with zinc dust.



OR Self-explanatory equation: 2Marks

Q26 Text Solution:

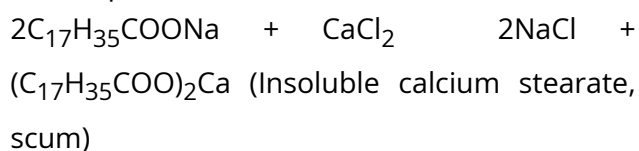
Acid chlorides is hydrogenated over catalyst, palladium on barium sulphate to give aldehyde.
 Rosenmund reduction.

Q27 Text Solution:

- (a) Food preservative
 (b) Antioxidant.

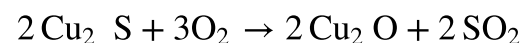
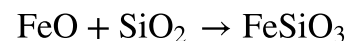
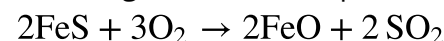
Q28 Text Solution:

Hard water contains calcium and magnesium ions. When soaps are dissolved in hard water, these ions displace sodium or potassium from their salts and form insoluble calcium or magnesium salts of fatty acids. These insoluble salts separate as scum.

**Q29 Text Solution:**

Copper matte is charged into silica lined convertor. Some silica is added and hot air blast is blown to convert the FeS to FeO and Cu₂S/ Cu₂O to the metallic copper.

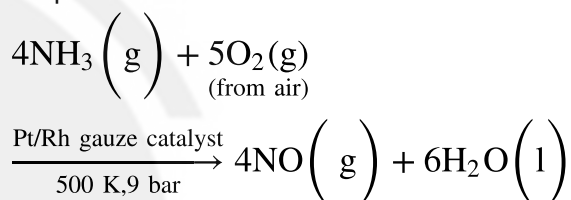
Following reactions take place:



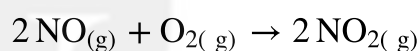
The solidified copper obtained has blistered appearance due to the evolution of SO₂ and so it is called blister copper

Q30 Text Solution:

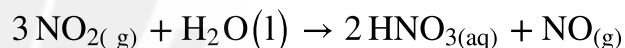
Step-1: Oxidation of ammonia to nitric oxide :



Step-2: Oxidation of NO to NO₂ :



Step-3: Formation of nitric acid :

**Q31 Text Solution:**

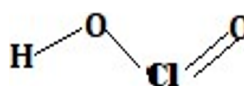
- (a) PbSO₄
 (b) 2Mn²⁺
 (c) 12C

Q32 Text Solution:

(a) By oxidation of hydrogen chloride gas by atmospheric oxygen in the presence of CuCl₂ at 723K.



(b) Structure of Chlorous acid is



Q33 Text Solution:

(a) (1) Due to variable (multiple) oxidation states.

(2) Large surface area for adsorption of reactant.

(3) Formation of intermediate compounds.

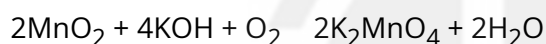
(4) due to their ability to form complexes

(b) For this question, definition for mischmetal is not available in NCERT textbook. Only its compositions, applications are given. Award one Mark if this question is attempted.

Q34 Text Solution:

KMnO_4 is manufactured from pyrolusite (MnO_2)

Step – I: Pyrolusite is powdered and fused with KOH in presence of KNO_3 as an oxidizing agent to form potassium manganate.



Step – II: The potassium manganate undergoes disproportionation in acidic or neutral medium to give permanganate.



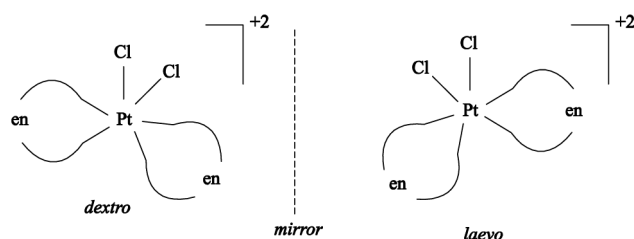
The purple solution so obtained is concentrated to get dark purple crystals of KMnO_4 .

Permanganate ion is also obtained by electrolytic oxidation of manganate ion in alkaline medium.

Q35 Text Solution:

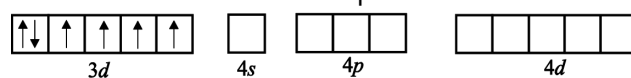
(a) $\text{cis} - [\text{PtCl}_2(\text{en})_2]^{2+}$ is optically active

(b) Structure of its enantiomers

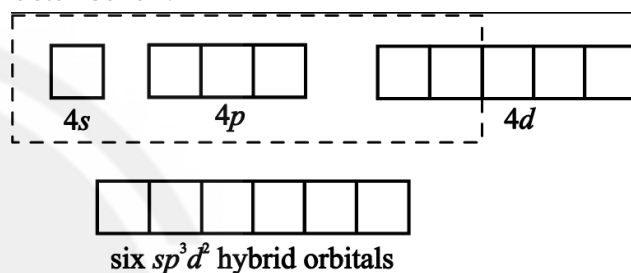
**Q36 Text Solution:**

In this complex, the oxidation state of Co is +3. The electronic configuration of cobalt in +3 oxidation state is $[\text{Ar}]3d^64s^0$.

Orbitals of Co^{3+} ion are represented as



Since F^- ion provides a weak ligand field, one 4s, three 4p and two outer 4d-orbitals hybridised to yield six sp^3d^2 hybrid orbitals pointing towards the six corners of an octahedron.



These six hybridised orbitals of Co^{3+} overlap with orbitals of six F^- ligands and six pairs of electrons donated by six F^- to form six coordinate bonds.



Six sp^3d^2 hybrid orbitals with six pairs of electrons from ligands.

Thus, the complex has octahedral geometry.

This complex is paramagnetic because of the presence of unpaired electrons.

Q37 Text Solution:

(a) The number of atoms per unit cell in fcc structures is four. Each atom is considered as one sphere.

Let the edge length of the unit cell = a

Radius of the sphere = r

Length of the face diagonal = b

$$\text{In } \triangle ABC, AC^2 = BC^2 + AB^2$$



$$b^2 = a^2 + a^2$$

$$b = \sqrt{2}a$$

$$\text{But } b = 4r$$

$$\therefore \sqrt{2} \cdot a = 4r$$

$$a = \frac{4r}{\sqrt{2}} = 2\sqrt{2}r$$

$$\text{Volume of one sphere} = \frac{4}{3}\pi r^3$$

Since FCC lattice contains 4 atoms (spheres) per unit cell,

$$\text{The volume of four spheres in fcc} = 4 \times \frac{4}{3}\pi r^3 = \frac{16}{3}\pi r^3$$

$$\text{The total volume of the unit cell} = a^3 = (2\sqrt{2}r)^3$$

$$\text{Packing efficiency} = \frac{\text{Volume of four spheres in unit cell}}{\text{Total volume of the unit cell}} \times 100$$

$$\text{Packing efficiency} = \frac{\frac{16}{3}\pi r^3}{(2\sqrt{2}r)^3} \times 100 = 74\%$$

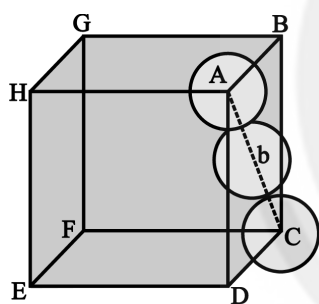


Fig. Unit cell of fcc

$$(b) d = \frac{z \times M}{a^3 \times N_A}$$

$$= \frac{2 \times 39}{(542 \times 10^{-10})^3 \times 6.022 \times 10^{23}} = \frac{78}{95.87} = 0.813 \text{ g cm}^{-3}$$

OR

$$= \frac{2 \times 39 \times 10^{-3}}{(542 \times 10^{-12})^3 \times 6.022 \times 10^{23}} = \frac{78 \times 10^{-3}}{95.87 \times 10^{-6}} = 813 \text{ kg m}^{-3}$$

Q38 Text Solution:

$$(a) M_2 = \frac{w_2 RT}{\pi v}$$

$$\pi = 3.1 \times 10^{-4} \text{ bar}, \quad V = 450 \text{ cm}^3 = 0.450 \text{ L}$$

$$T = 310 \text{ K}, \quad R = 0.083 \text{ L bar K}^{-1} \text{ mol}^{-1} \quad w_2 = 1.0 \text{ g}$$

$$M_2 = \frac{1.0 \times 0.083 \times 310}{3.1 \times 10^{-4} \times 0.450}$$

$$= \frac{25.73 \times 10^4}{1.395}$$

$$= 1,84,444 \text{ g mol}^{-1}$$

(b) Statement of Raoult's law of Relative lowering of vapour pressure is not available in NCERT textbook.

This question may be treated as out of the NCERT text book syllabus. Award two marks if this question is attempted.

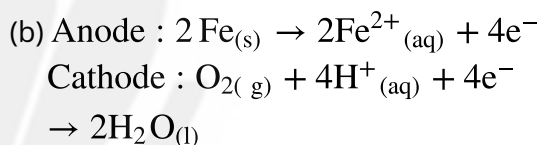
Q39 Text Solution:

$$(a) E_{\text{cell}}^0 = E_{\text{Cathode}}^0 - E_{\text{Anode}}^0 = E_{\text{Ag}}^0 - E_{\text{Zn}}^0$$

$$= 0.80 - (0.76) = 1.56 \text{ V}$$

$$\Delta G^0 = -nFE_{\text{Cell}}^0 = -2 \times 96500 \times 1.56 = -301080 \text{ J mol}^{-1}$$

$$\text{or } -301.080 \text{ kJ mol}^{-1}$$



Q40 Text Solution:

(a) Consider a first order reaction, $R \rightarrow P$

A first order reaction is one in which the rate is directly proportional to first power of the reactant concentration.

Therefore, according to rate law,

$$\text{Rate} \propto [R]^1$$

$$\text{Rate} = k[R]^1 \dots (1)$$

Where k is rate constant or velocity constant

$$\text{But, Rate} = -\frac{d[R]}{dt}$$

$$-\frac{d[R]}{dt} = k[R] \dots (2)$$

Rearrange the equation (2), we get

$$\frac{d[R]}{[R]} = -k dt \dots (3)$$

$$\text{Integrate equation (3)} \int \frac{1}{[R]} d[R] = -k \int dt$$

$$\ln[R] = -kt + I \dots (4)$$



When $t = 0$, $[R] = [R]_0$ where $[R]_0$ is the initial concentration of reactant R.

$$\ln [R]_0 = -k \times 0 + I$$

where I is called integration constant

$$I = \ln [R]_0$$

Substituting the value of I in equation (4) we get,

$$\ln [R] = -kt + \ln [R]_0$$

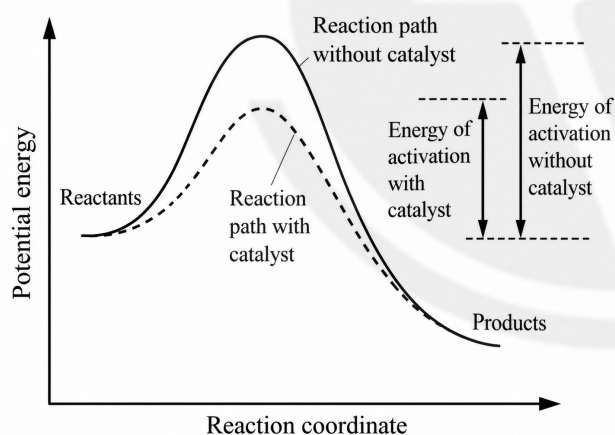
$$k \ln [R]_0 - \ln [R]$$

$$kt = \ln \frac{[R]_0}{[R]}$$

$$kt = 2.303 \log_{10} \frac{[R]_0}{[R]}$$

$$k = \frac{2.303}{t} \log_{10} \frac{[R]_0}{[R]}$$

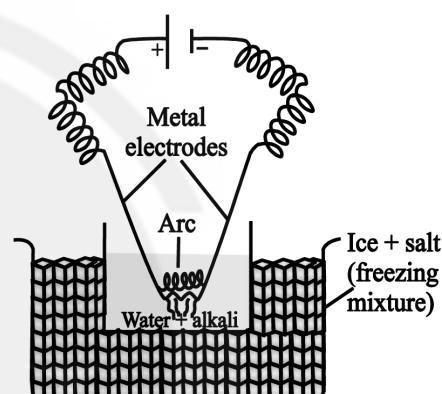
(b)



Q41 Text Solution:

(a) Bredig's arc method involves dispersion as well as condensation & used to prepare metal sols like gold sol, silver sol, platinum sol, etc.

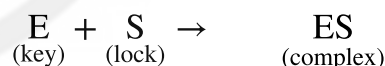
Two gold rods are dipped in ice cold water containing little KOH (alkali is added to stabilise gold sol). The vessel is kept in freezing mixture. An electric arc is struck between two gold rods. Heat produced by the spark causes a small amount of gold to vapourize. The vapours of gold condense suddenly to form particles of colloidal size of colloidal gold or gold sol. (figure 1 one mark)



(b) Step 1: Binding of enzyme to substrate to form an activated complex.



OR

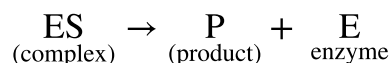


Lock and key model

Step 2: Decomposition of the activated complex to form product.



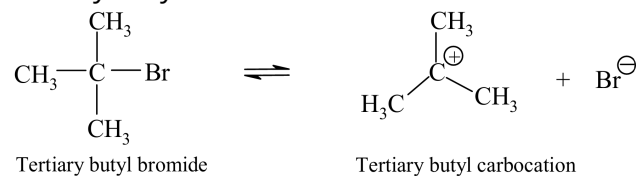
OR



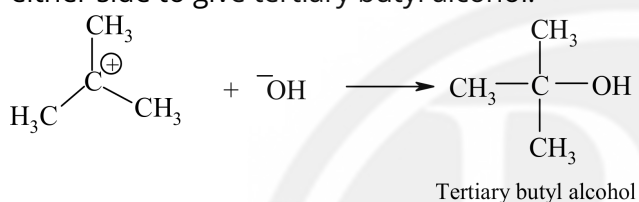
Q42 Text Solution:

(a) S_N1 mechanism involves two steps. It follows first order kinetics.

I Step: (slow step): Tertiary butyl bromide ionizes slowly to give sp^2 hybridised planar tertiary butyl carbocation and bromide ion.



II Step: (Fast step)→The nucleophile OH^- from aqueous NaOH attacks planar carbocation on either side to give tertiary butyl alcohol.



The product tertiary butyl alcohol obtained is a racemic mixture.

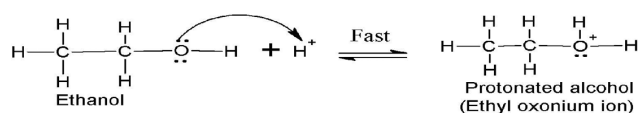
(b) (i) Resonance effect or C – X bond acquires a partial double bond character due to resonance

(ii) Due to sp^2 hybridization of carbon atom in C – X bond

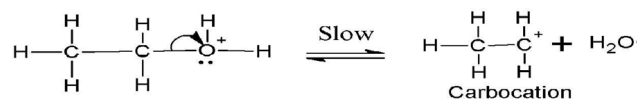
(iii) Instability of phenyl cation

Q43 Text Solution:

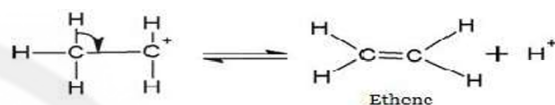
(a) Step1: Formation of protonated alcohol:



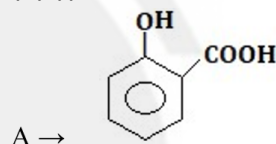
Step2: Formation of carbocation:



Step3: Elimination of proton:



(b) (i)



OR

o - hydroxybenzoic acid OR Salicylic acid

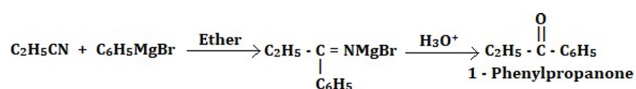
(ii) B $(\text{CH}_3)_3\text{C-I}$ OR t-butyl iodide

OR 2-iodo-2-methylpropane

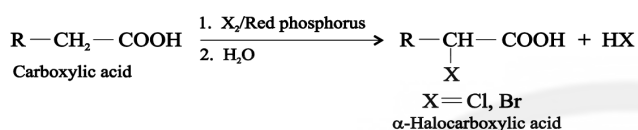


Q44 Text Solution:

(a) Treatment of nitrile with Grignard reagent followed by hydrolysis gives a ketone.



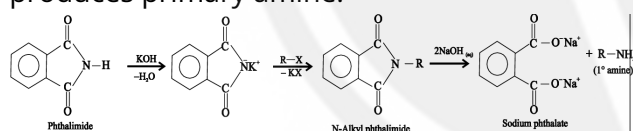
(b) Carboxylic acids having an α -hydrogen on treated with chlorine or bromine in the presence of red Phosphorus gives α -halocarboxylic acid.



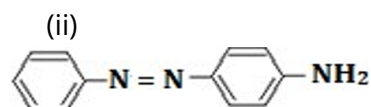
(c) Dry hydrogen chloride protonates the carbonyl oxygen & increase the electrophilicity of the carbonyl carbon.

Q45 Text Solution:

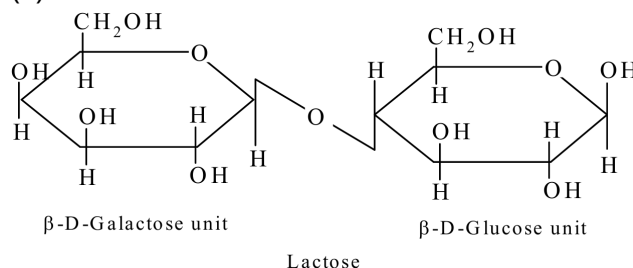
(a) Phthalimide on treatment with ethanolic potassium hydroxide (KOH) form potassium salt of phthalimide which on heating with alkyl halides followed by alkaline hydrolysis produces primary amine.



(b) (i) $\text{C}_6\text{H}_5\text{N}_2^+\text{Cl}^-$ or $\text{C}_6\text{H}_5\text{N}_2\text{Cl}$

**Q46 Text Solution:**

(a)



Loss of biological activity of protein by heating or change in temperature or pH is called denaturation of protein.

OR

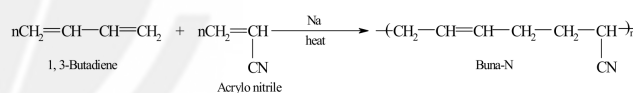
(b) Co-agulation of protein is called denaturation of protein.

Primary (1^0) structure remains intact.

(c) Deoxyribose OR β -D-2-deoxyribose;

Q47 Text Solution:

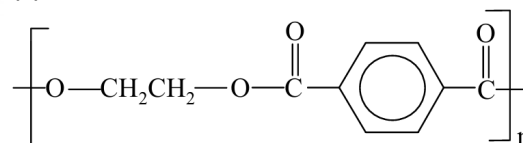
(a) Buna - N manufactured by the co-polymerisation of 1, 3-butadiene and acrylonitrile in presence of peroxide or sodium catalyst.



(b) Glycine

Aminocaproic acid

(c)



Android App

iOS App

PW Website